THE RELATIONSHIP BETWEEN THE COMBUSTION HEAT OF ORGANIC COMPOUNDS AND THE NUMBER OF GIVEN ELECTRONS

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Until now, only one letter has been written about this subject. According to the result of this letter, the combustion heat of organic compounds depends on the oxidation step of carbon. The combustion heats of alkanes, alkenes and alcohols are investigated. In our study, the combustion heats of aromatic compounds were investigated. Firstly, the combustion reactions of aromatic compounds were written and balanced according to oxidation number method. Then, the number of given electrons for one molecule was determined. As a result of the ratio (the combustion heats of organic compounds /the number of given electrons for one molecule), I reached a constant result. According to the result of my study, when we multiply the number of given electrons for one molecule and this obtained constant, we get the combustion heat of organic compounds. The constant for the combustion heat of aromatic compounds is -26,221 kcal/mol·e⁻.